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You are given three clear, colorless,

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unknown solutions, and red and blue litmus papers. Describe the steps you might take to correctly identify each solution as acidic, basic, or neutral. 29. Describe how solutions of acids and bases differ in terms of particles. How do the differences relate to the properties of acids and bases? 30.

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HSAB concept is an initialism for "hard and soft (Lewis) acids and bases". Also known as the Pearson acid-base concept, HSAB is widely used in chemistry for explaining stability of compounds, reaction mechanisms and



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pathways. It assigns the terms 'hard' or 'soft', and 'acid' or 'base' to chemical species. 'Hard' applies to species which are small, have high charge states (the charge criterion ...

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smaller than the particles in a 6. A 7. D  
8. A 9. A 10. A 11. colloid 12. separate  
the ions from each other. dissolve 13.  
corrosive 14. hydrochloric 15. base 16.  
F, water and a salt 17. F, temperature  
18. T 19. F, bitter 20. F ...

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Acids, Bases, and Solutions ANSWER KEY  
Acids, Bases, and Solutions Describing  
Acids and Bases Review and Reinforce 1.  
Sour 2. Bitter 3. Corrosive to  
magnesium, zinc, and iron; eats them  
away and produces bubbles of hydrogen  
gas 4. Doesn't react with metals 5.

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Produces carbon dioxide 6. Doesn't react with carbonates 7. Red 8. Blue 9.

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Given acids or bases at the same concentration, demonstrate understanding of acid and base strength by:

1. Relating the strength of an acid or base to the extent to which it dissociates in water
2. Identifying all of the molecules and ions that are present in a given acid or base solution.
3. Comparing the relative concentrations of molecules

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and ...

## **Acid-Base Solutions - Acids | Bases | Equilibrium - PhET ...**

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throughout; a mixture in which one substance is dissolved into another. solvent.

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Chapter 7 Acids, Bases, and Solutions  
Calculating a Concentration To calculate the concentration of a solution, compare

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the amount of solute to the amount of solution and multiply by 100 percent. For example, if a solution contains 10 grams of solute dissolved in 100 grams of solution, then its concentration can be reported as 10 percent.

## **Chapter 7 Acids, Bases, and Solutions**

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Practice: Acids, bases, and pH. Sort by: Top Voted. Introduction to buffers. Acids, bases, and pH. Up Next. Acids, bases, and pH. Biology is brought to you with support from the Amgen Foundation. Biology is brought to you with support from the. Our mission is to provide a free, world-class education to anyone, anywhere.



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Chapter 7 Acids, Bases, and Solutions

Instructions: Complete the crossword puzzle. Use the clues to help identify the words. C O N C E N T R A T E D H O E I S  
C A L E Y L U L O D L T U N S A T U R A T  
E D R O R T A R O I A E L O G D L T S A T

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URATED INDICATOR IN ZV AS  
SUPERSATURATED

## **Chapter 7 Acids, Bases, and Solutions**

Acids that dissociate to a large extent  
are strong electrolytes and strong acids

Acids that dissociate to only a small  
extent are weak electrolytes and weak

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acids Strong vs Weak Acids Arrhenius, Bronsted-Lowry and Lewis Theory of Acid and Bases Weak Acid John E. McMurry, Robert C.Fay, Jill K. Robinson. Chemistry. 7 th edition. Pearson Education ...

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(b) Brønsted-Lowry's acid and base. In a new theory submitted in 1923 independently by Brønsted and Lowry, an acid is defined as a molecule or an ion which gives  $H^+$  and a molecule or ion that receives  $H^+$  from a partner is a base. A base is not only a molecule or an ion which gives  $OH^-$  but anything which receives  $H^+$ . Since the acid  $HA$  gives  $H^+$

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+ to water in an aqueous solution and generates ...

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Section 8.3 Properties of Acids and Bases (pages 240–245) This section describes the general properties of acids and bases. Reading Strategy (page 240)

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