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Extension Questions

15. One mole of any gas will occupy 22.4 L of volume at standard temperature and pressure (STP).

Consider this situation as it relates to the reaction in Model 1:  $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$

a. Calculate the volumes of reactants consumed and the volume of product made.

b. Record the ratio of  $N_2$  to  $H_2$  to  $NH_3$

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Answer Key. "The  
""Who is your optimal  
friend"" job interview  
problem is utilised  
sometimes as part  
from the behavioral  
assessment portion of  
the task job interview.  
The solution you give  
to this question  
furnishes the hiring  
manager having a

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sense as to how introspective you might be in terms of how other men and women view you, and nicely you map your personality characteristics into a succeed ecosystem.

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## Answers

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Activity Topic: ... show how you calculated your answer.  $21.0 \text{ mol A} * 5 \text{ mol C} = 52.5 \text{ mol C}$ . ... mole ratio between moles of given to moles of  $\hat{\text{€}}$

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KEY Write the empirical formula for the following compounds.

1)  $C_6H_6$  C 3 H 3 6)

$C_8H_{18}$  C 4 H 9 7)  $WO_2$

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## Answers

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### **Relative Mass And The Mole Answer Key**

Answer: 0.234 g SO<sub>2</sub>  
(1 mol SO<sub>2</sub> 64.07 g  
SO<sub>2</sub>) (1 mol O<sub>2</sub> 2 mol  
SO<sub>2</sub>) (22.4 L O<sub>2</sub> 1 mol  
O<sub>2</sub>) = 0.0409 L O<sub>2</sub> (In  
your calculator: 0.234  
÷ 64.07 ÷ 2 × 22.4 =)  
Molar Mass Mole-Mole  
Ratio Molar Mass

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## Answers

Grams of Given Moles  
of Given Moles of  
Unknown Grams of  
Unknown Molar Mass  
Mole-Mole Ratio Molar  
Volume gas @ STP  
Grams of Given Moles  
of Given Moles of  
Unknown

## **Chapter 13** **Stoichiometry**

A dozen is twelve things, and a mole is an Avogadro's number ( $6.022 \times 10^{23}$ ) of things. The The

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difference is that a dozen things (12) is an exact number, but a mole of things ( $6.022 \times 10^{23}$ )

## **Moles & Stoichiometry Answers Key Questions & Exercises**

The answer is the unit called the mole. This activity will start by considering two egg farmers (a chicken farmer and a quail

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## Answers

farmer). They produce such large numbers of eggs that they can't count them all individually, so they count in dozens of eggs in some cases, while in other cases they use mass.

### **Relative\_Mass\_and\_the\_Mole\_POGIL.docx - Relative Mass and**

...

The ratio of masses for a mole of oxygen and a mole of sulfur is 1:2, so

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## Answers

the ratio of number of atoms in 16- and 32-1b samples will still be 1: 1.

II. In the front of the room, there is a bottle that contains a 32.00 g sample of sulfur. This is 1 mole of sulfur. Estimate how many atoms are in the bottle. Your group must reach consensus.



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