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There is a simple set of rules for these acids.

Any polyatomic ion with the suffix “-ate” uses the suffix “-ic” as an acid. So, HNO_3 will be nitric acid. When you have a polyatomic ion with one more oxygen than the “-ate” ion, then your acid will

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have the prefix “per-” and the suffix “-ic.” For example, the chlorate ion is ClO_3^- .

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important substances
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life we encounter acids

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such as vinegar and lemon juice and bases such as ammonia and lye in drain cleaners. Many important processes in chemistry and biology involve reactions of acids with bases . . .

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perchloric acid: HCOOH

formic acid: H_3PO_4

phosphoric acid: HCl

(aq) hydrochloric acid:

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H₃BO₃ boric acid: H
2SO₄ sulfuric acid:
HNO₂ nitrous acid: HI
(aq) hydroiodic acid:
CH₃COOH acetic acid:
HF (aq) hydrofluoric
acid: H₃PO₃
phosphorous acid: HCN
(aq) hydrocyanic acid:
HClO₃ chloric acid: H₂
CO₃ carbonic acid: H₂
SO₃ sulfurous acid:
HClO₂ chlorous acid:
HNO₃ nitric acid: HBr
(aq) hydrobromic acid

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13. All acid—base reactions have two conjugate acid—base pairs. One conjugate acid—base pair in the reaction in Model 3 is $\text{H}_2\text{O}/\text{H}_3\text{O}^+$. List the other acid—base pair in the reaction. 14. Why is HCO_3^- considered the "acid" part of the pair in the reaction in Model 3? 15. Is CO_3^{2-} considered the "base" part of the

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Model 3? 16.

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