

## Problem Solving Ph Chemfile Answers

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Holt ChemFile: Problem-Solving Workbook 260 pH Name Class Date Problem Solving continued The negative logarithm of 10<sup>-14</sup> is 14. Substitute this value, and rearrange. pOH = 14 - pH = 14 - 1.2 = 12.8. General Plan for Solving pH Problems Mass of solute Convert using the molar mass of the solute. This step works only for strong acids or bases. Multiply the concentration of the solution by

### Skills Worksheet Problem Solving

Problem : What is the pH of a 0.001 M solution of H<sub>2</sub>SO<sub>4</sub>? HSO<sub>4</sub><sup>-</sup> has a pK<sub>a</sub> of 1.2 × 10<sup>-2</sup>. To solve this problem, you must first note that sulfuric acid's first deprotonation is as a strong acid, so we have a concentration of 0.001 M H<sup>+</sup> to start and 0.001 M hydrogen sulfate. Because hydrogen sulfate is a weak acid, this problem becomes very similar to the last one (see ).

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Holt ChemFile: Problem-Solving Workbook 261 pH Name Class Date Problem Solving continued Sample Problem 1 A HCl solution has a concentration of 0.0050 M. Calculate  $[\text{OH}^-]$  and  $[\text{H}_3\text{O}^+]$  for this solution. HCl is a strong acid, so assume it is 100% ion-ized.

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20/04/2016 · Holt ChemFile: Problem-Solving Workbook 259 pH pH In 1909, Danish biochemist S P L Sørensen introduced a system in which acidity was expressed as the negative logarithm of the H concentration In this way, the acidity of a solution having H

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Bookmark File PDF Holt Chemfile Problem Solving Workbook Ph needed 0.855 g Cl<sub>2</sub> 1 mol Cl<sub>2</sub> 70.90 g Cl<sub>2</sub> 0.0121 mol Cl<sub>2</sub> 3.205 g KBr Skills Worksheet Problem Solving

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Holt ChemFile: Problem-Solving Workbook 57 Mole Concept Name Class Date Problem Solving continued Is the answer reasonable? Yes; 2 g of boron is about 1/5 of the molar mass of

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boron. Therefore, 2.00 g boron will contain about  $1/5$  of an Avogadro's constant of atoms. Practice 1. Calculate the number of atoms in each of the following masses:

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