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4.8. The rate of the chemical reaction doubles for an increase of 10 K in absolute temperature from 298 K.

Calculate  $E_a$ . Ans. 4.9. The activation energy for the reaction,  $2 \text{HI}(\text{g}) \rightarrow \text{H}_2 + \text{I}_2(\text{g})$  is  $209.5 \text{ kJ mol}^{-1}$  at 581

K. Calculate the fraction of molecules of reactants having energy equal to or

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greater than activation energy?

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molecules in the gas phase

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